



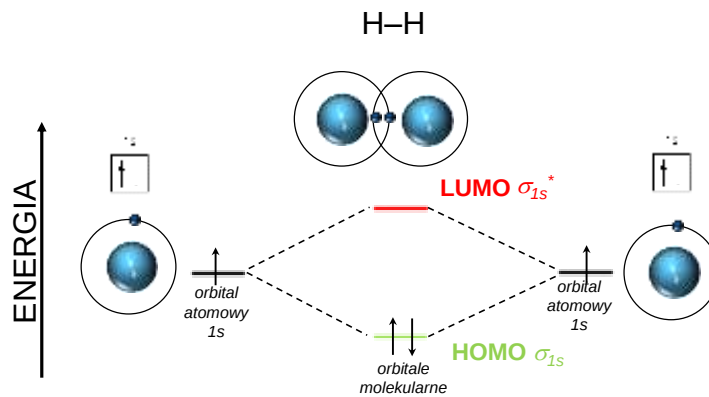
<http://scholaris.pl/>

<http://www.youtube.com/>

<http://wikipedia.com/>

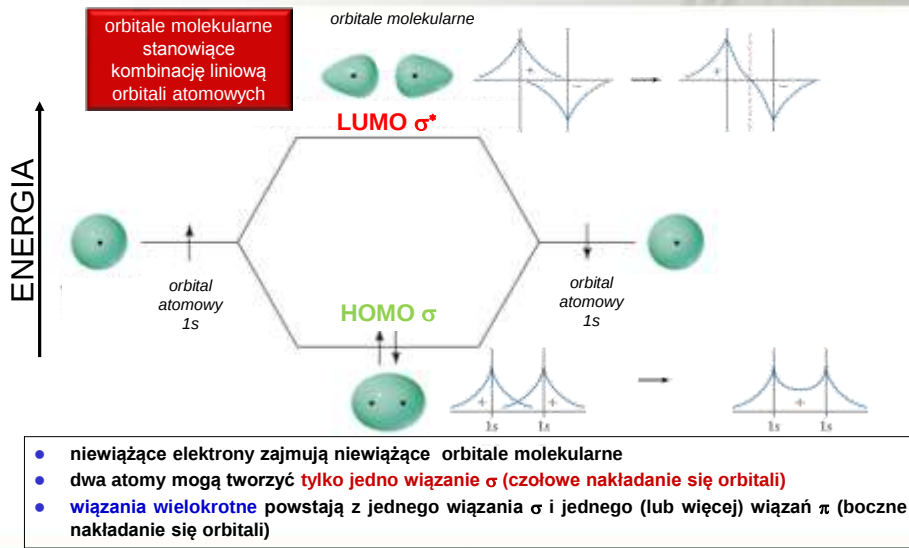


Orbitale molekularne H_2

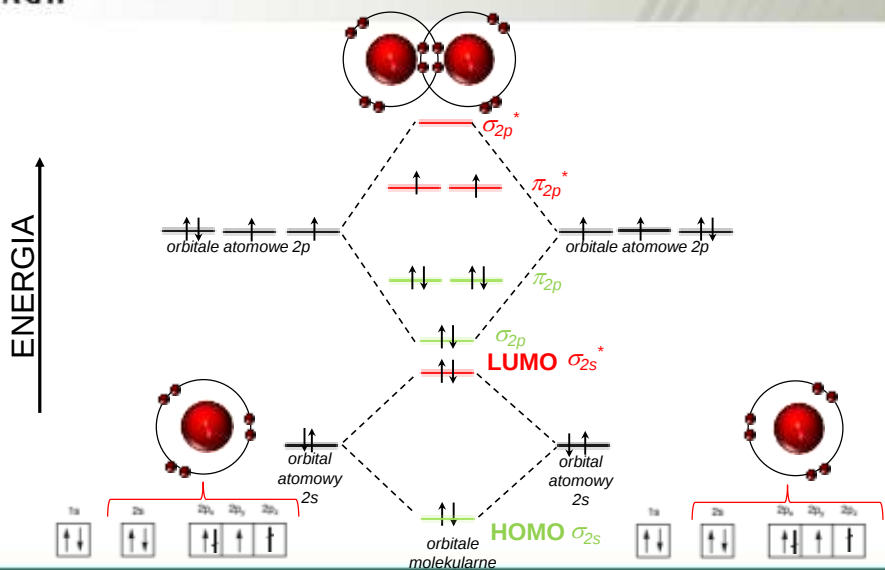




Kształt orbitali molekularnych (nakładanie s-s)

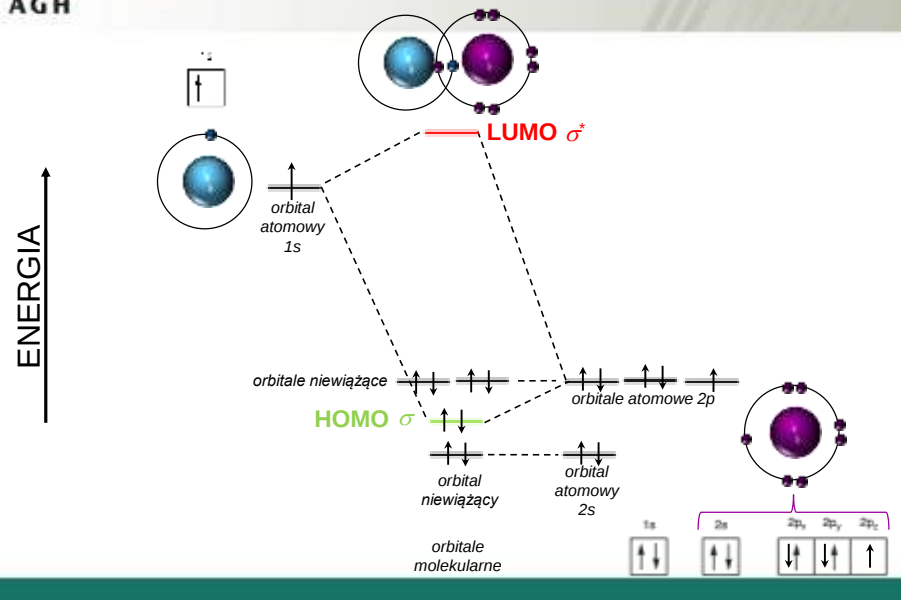


Orbitale molekularne O₂

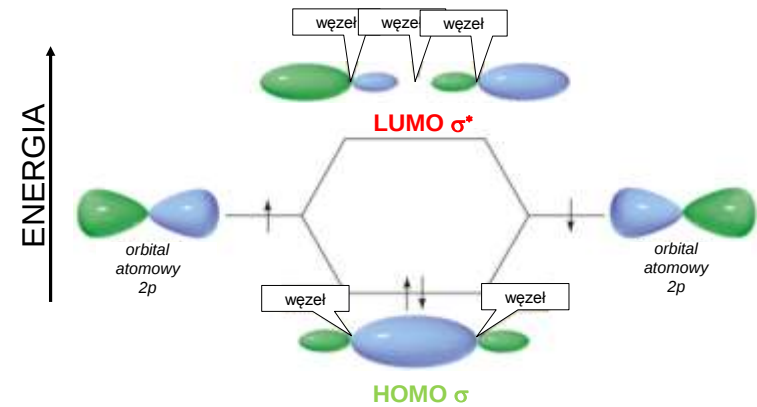




Orbitale molekularne HF

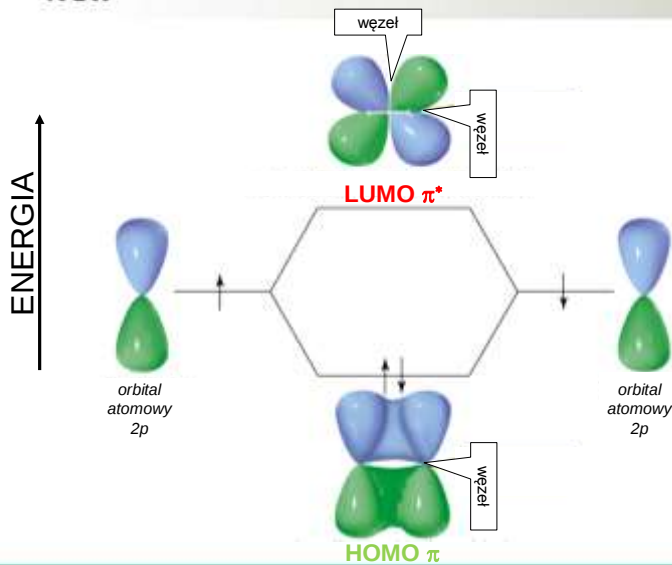


Kształt orbitali molekularnych (nakładanie p-p)

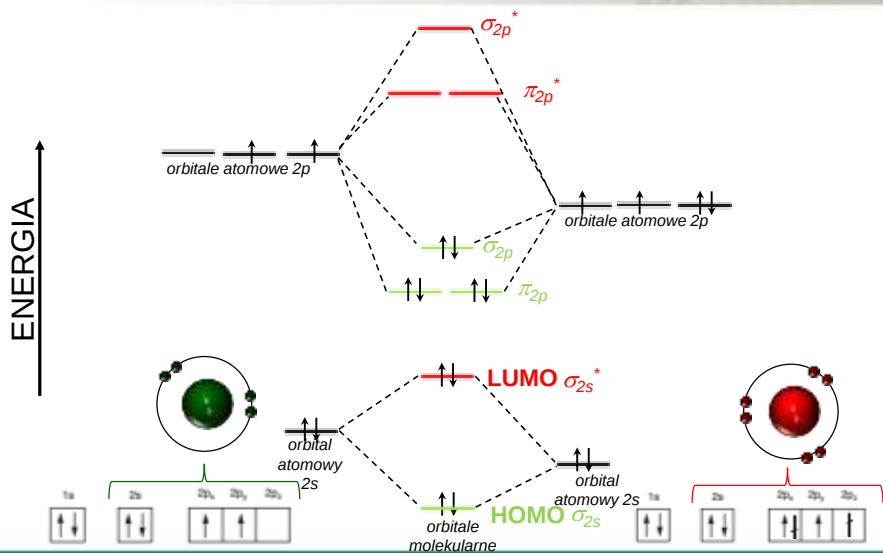




Kształt orbitali molekularnych (nakładanie p-p)



Orbitale molekularne CO





Orbitale molekularne CO

Covalent bond

Valence Bond Theory

Carbon Monoxide (orbitals)

www.quimica3d.com
September, 2009



Orbitale molekularne CO

Molecular Orbital Theory

Covalent Bond

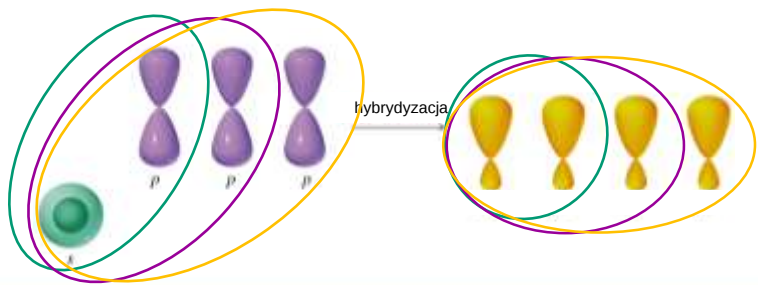
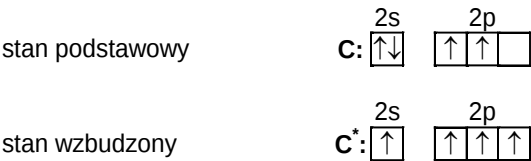
Carbon Monoxide

www.quimica3d.com
June, 2009
Updated September, 2010

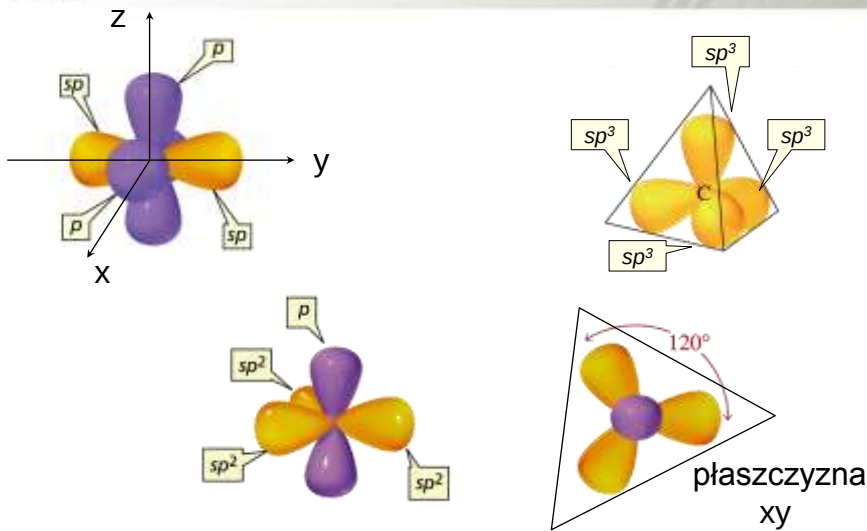




Hybrydyzacja

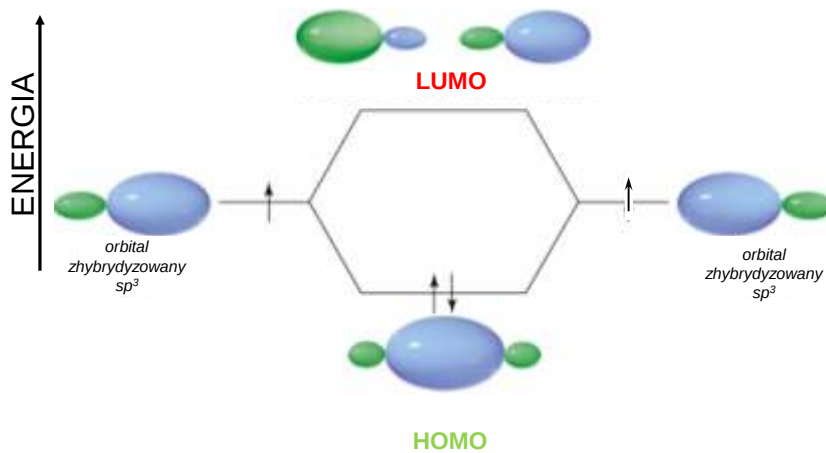


Hybrydyzacja

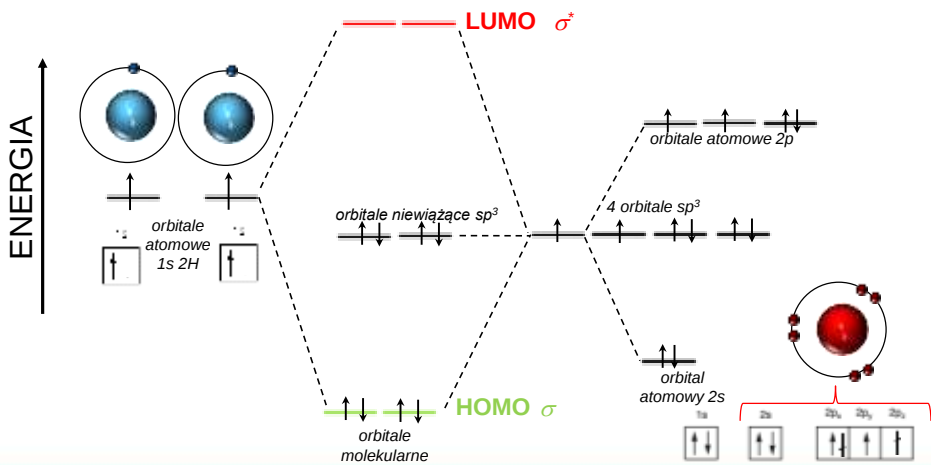




Orbitale molekularne

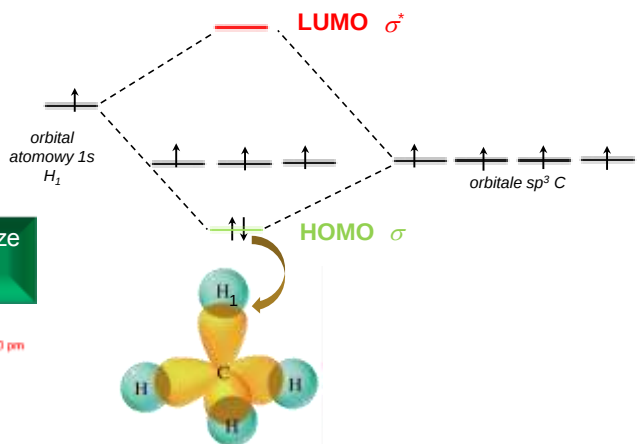
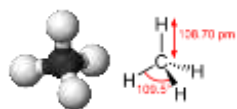


Orbitale molekularne H_2O

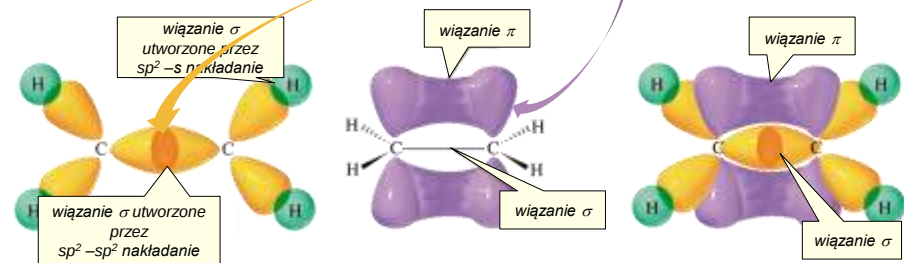
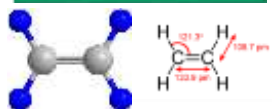




wiązanie pojedyncze
metan, CH_4

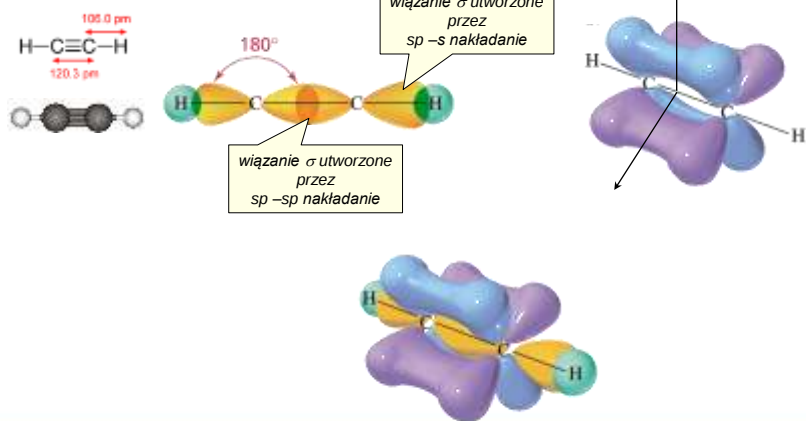


wiązanie podwójne
eten, $\text{CH}_2=\text{CH}_2$





wiązanie potrójne
etyń, $\text{CH}\equiv\text{CH}$



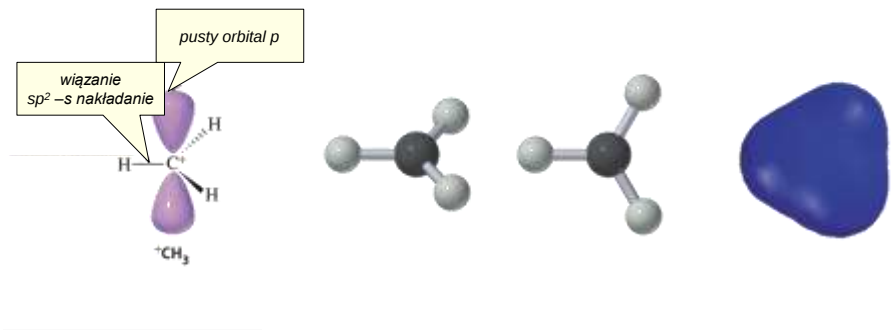
Siła wiązań

wiązanie	typ	długość	energia		molekuła
C-C	sp^3-sp^3	1.54 Å	Kcal 88	(KJ) (368)	CH_3-CH_3
C=C	sp^2-sp^2 and p - p	1.34 Å	145	(607)	$\text{CH}_2=\text{CH}_2$
C≡C	$\text{sp} - \text{sp}$	1.21 Å	198	(828)	$\text{HC}\equiv\text{CH}$

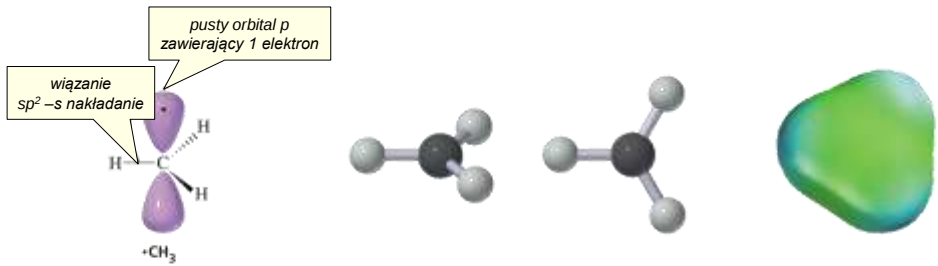
wzrost
charakteru „s”



Kation metylu

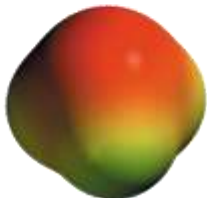
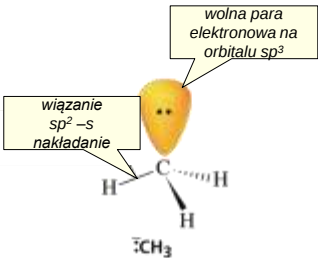


Rodnik metylu

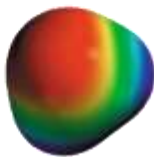
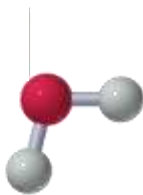
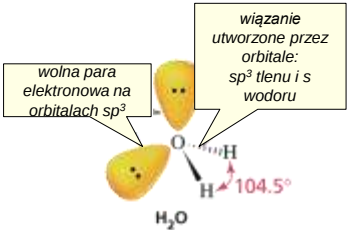




Anion metylu



Woda





Amoniak i jon amoniowy

